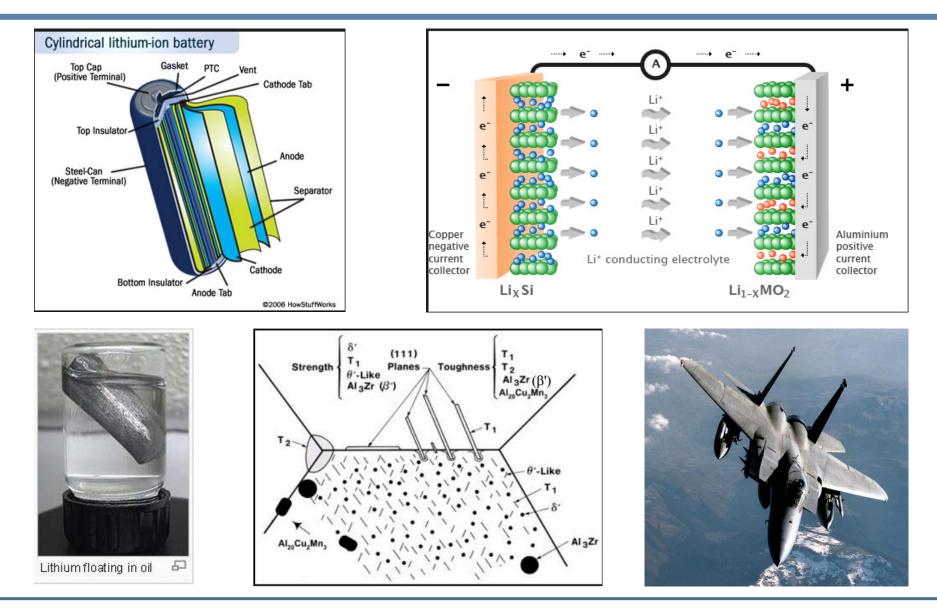


Detecting Li with EDS

October 2014

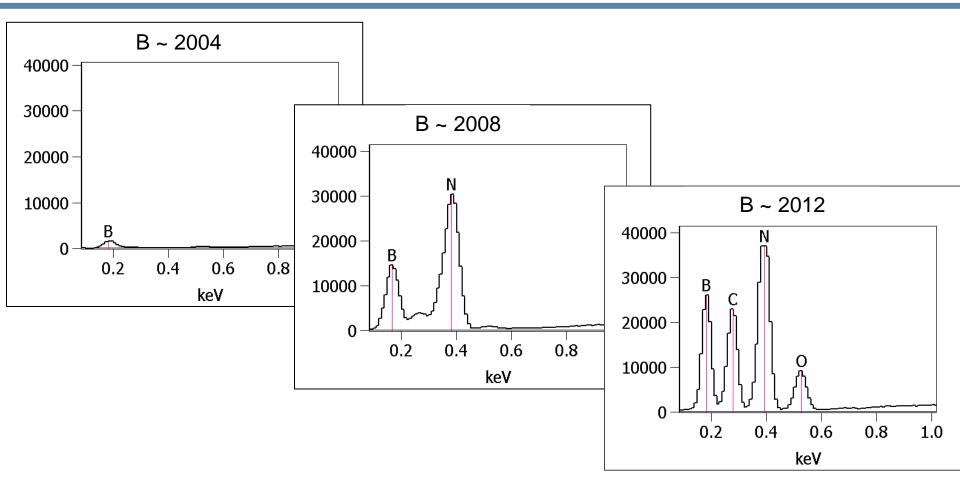
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Why detect Li x-rays? Applications





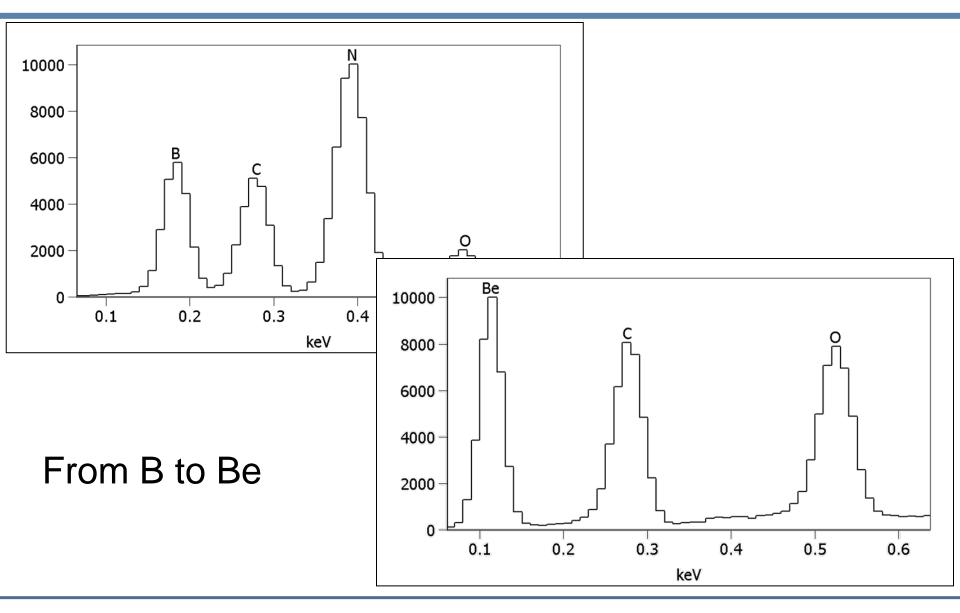
Why detect Li X-rays? Detector Technology



Light element sensitivity represents a continued evolution in detector technology.



Why detect Li X-rays? Detector Technology

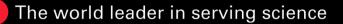






Challenge # 1

Building an EDS that detects Li x-rays



Building an EDS that detects Li x-rays

Spectral performance

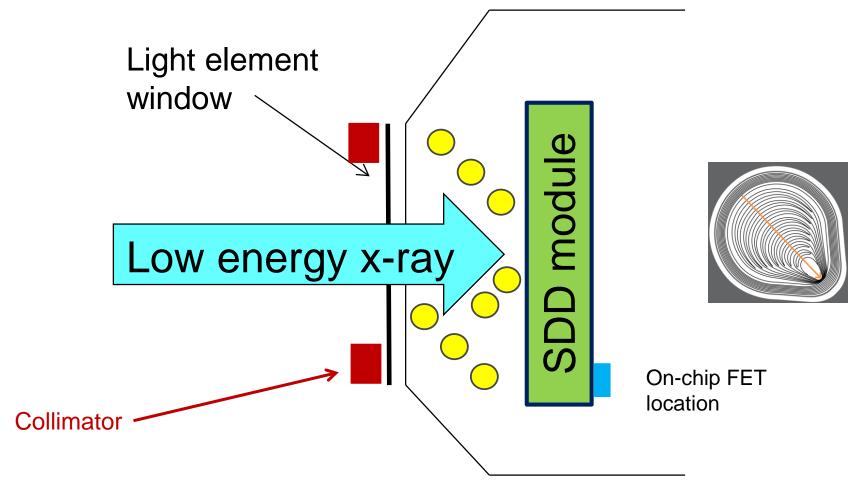
- Zero width < 30 eV
- Discriminate a 52 eV energy event
- Peak to background for trace detection

Windowless

• Eliminate absorption of Li x-rays in the window



Building an EDS that detects Li x-rays

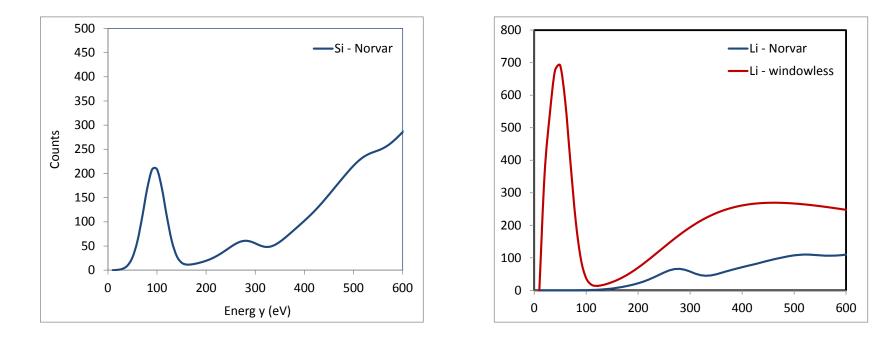


Low energy x-rays have a difficult time getting to the SDD module



The challenge of the window

Simulations

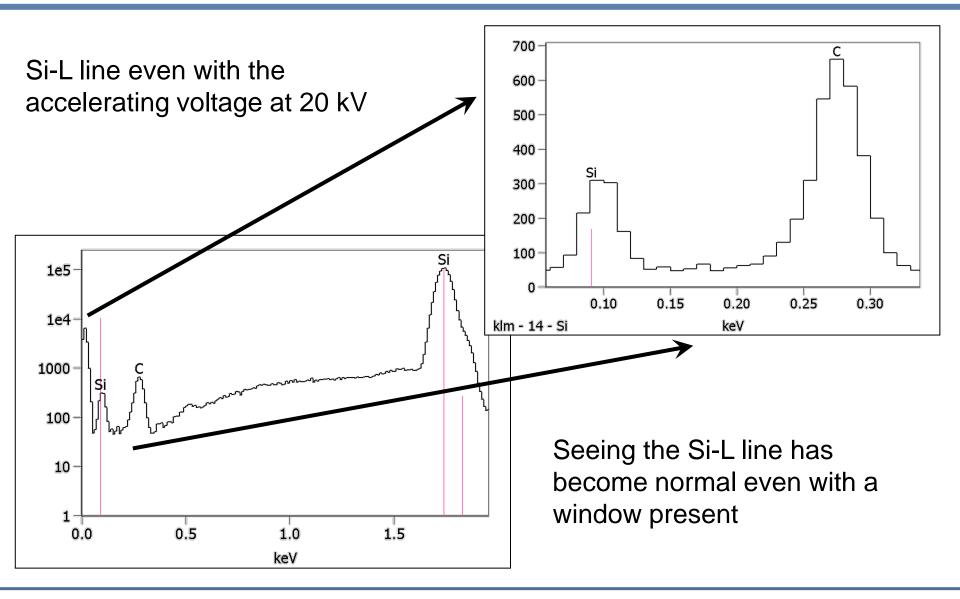


Si-I line with window

Li-k line with and w/o window

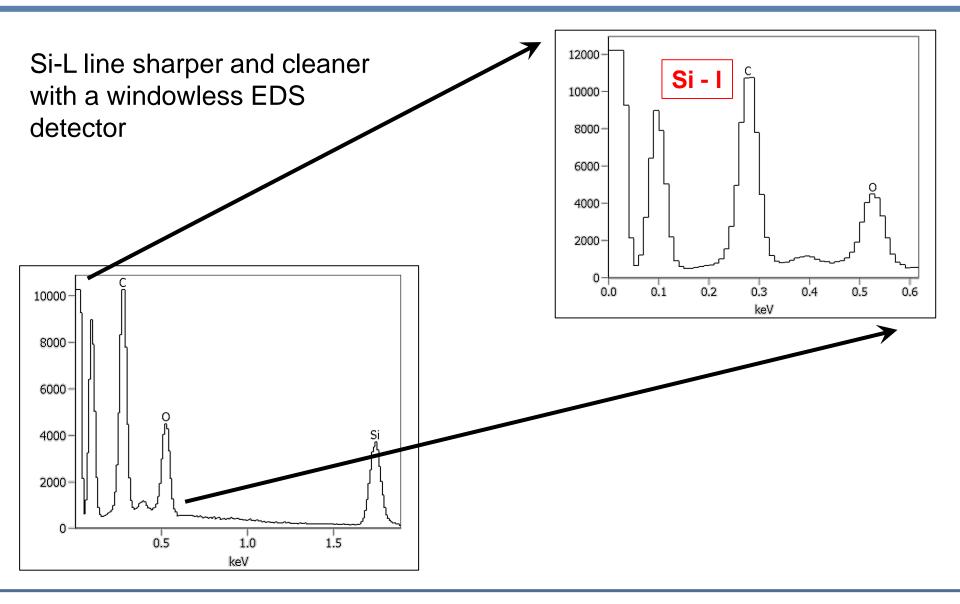


Actual data: Si-L lines (detector w/window)



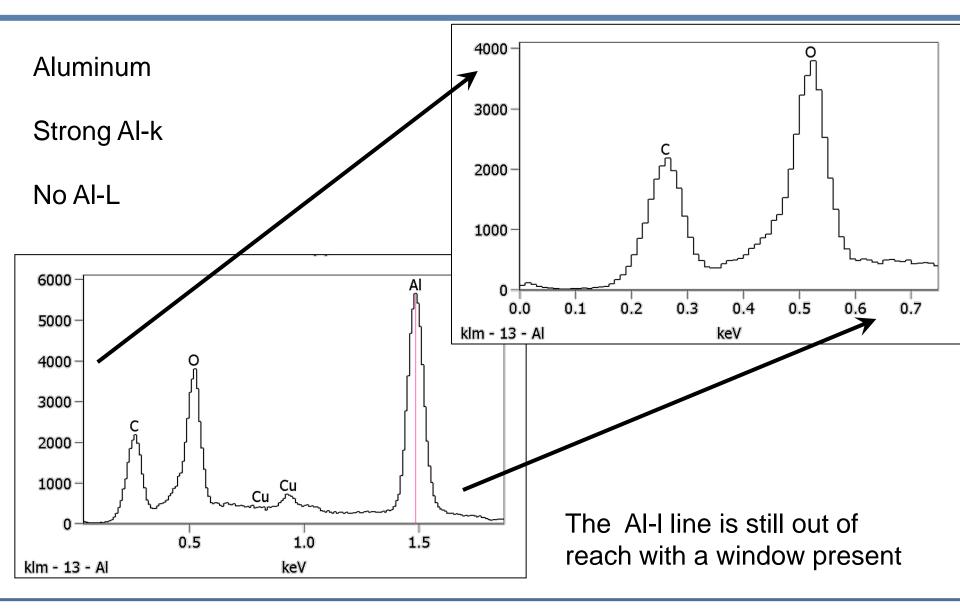


Actual data: Si-L lines (windowless)



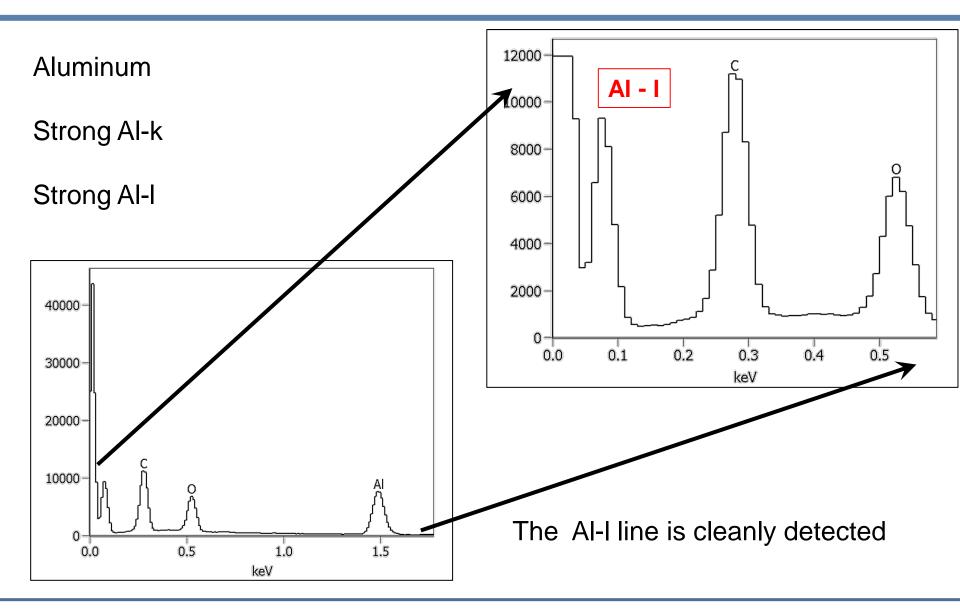


Actual data: Al lines (detector w/window)



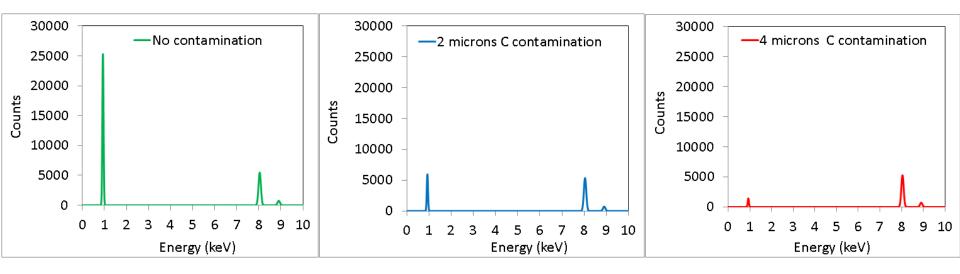


Actual data: Al lines (windowless)





Impact of module contamination Simulations

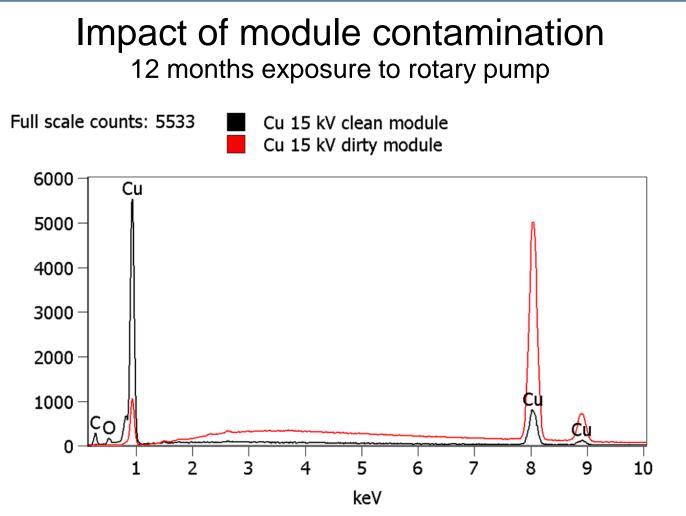


Copper k (~ 8 keV) and I (~ 1 keV) lines

Module is -25° C cold Window is insulated and therefore relatively warm



The challenge of no window



Module is -25° C cold Window is insulated and therefore relatively warm



Detector requirements for Li detection

Spectral performance

- Zero width < 30 eV
- Discriminate a 52 eV energy event
- Peak to background for trace detection

Safe Windowless operation

- Eliminate absorption of Li x-rays in the window
- Exposed (-25° C cold) SDD module must be protected when not in use with interlocks against chamber vent at all levels.
 - Pressure sensor, software time-outs, etc.
 - Shutter when not in use
- SEM must be a dry-pumped system. Oil in the chamber will contaminate the module and ruin the Li detection capability.







Getting Li x-rays out of the sample

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Sample requirements

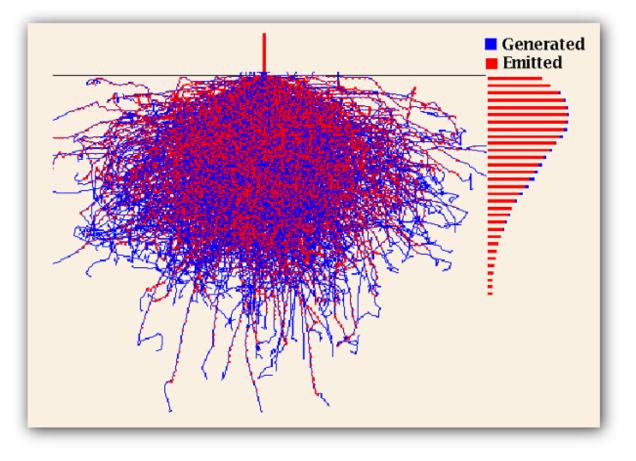
- Must get x-rays out of the samples to measure them
 - X-ray absorption within the sample
 - Heavy elements are really effective at absorbing low energy x-rays
- Li is highly reactive. When exposed to air, it oxidizes almost immediately.

 The escape depth of Li is only 40 nm – 50 nm. As the surface oxidizes, the Li x-rays are increasingly absorbed in the oxidized layer.



Absorption example

 Fe Kα X-Ray emission from an FeB sample

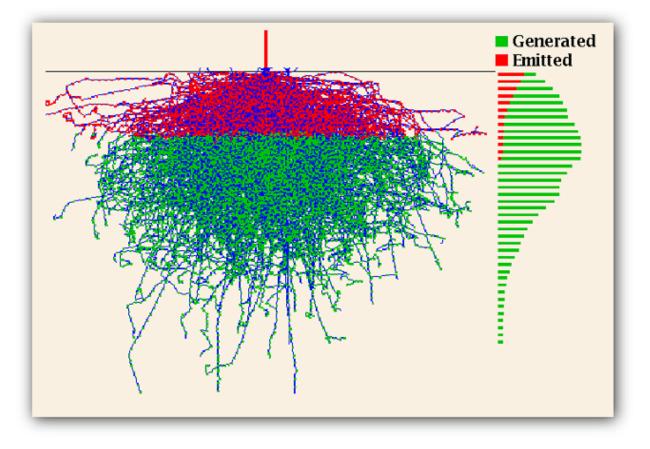




Photon - Sample Interactions

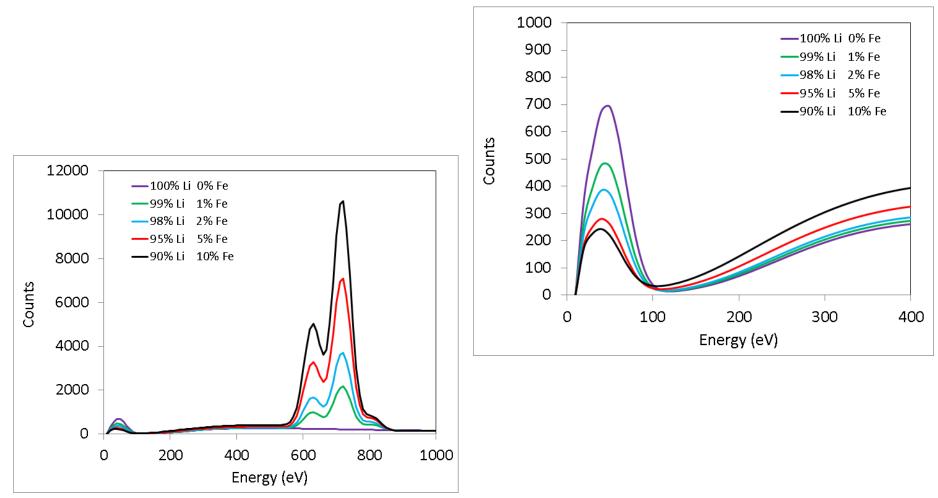
Absorption example

 B Kα X-Ray emission from an FeB sample





Absorption of fluoresced x-rays



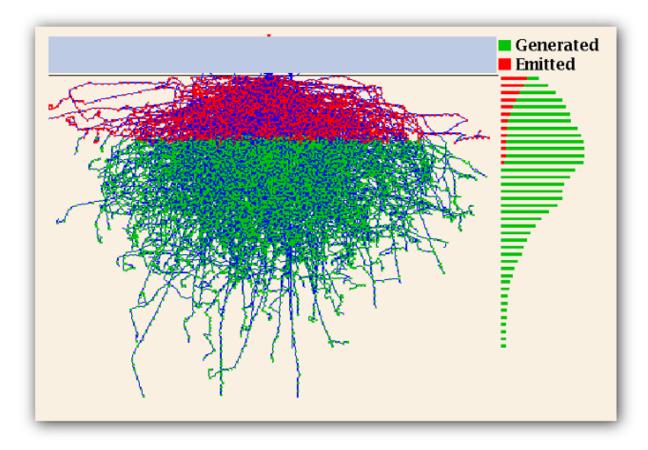
- Relative peak heights of Fe and Li as a small amount of Fe is added
- As we add even a small amount of a heavier element, the Li is strongly absorbed



Effect of a Thin layer on surface

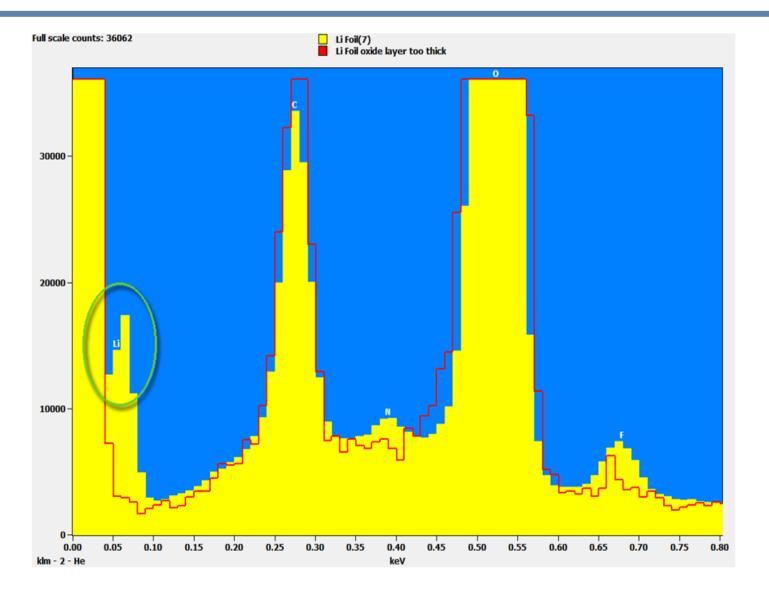
Absorption example

 As a thin layer grows on the surface, low energy x-rays are absorbed



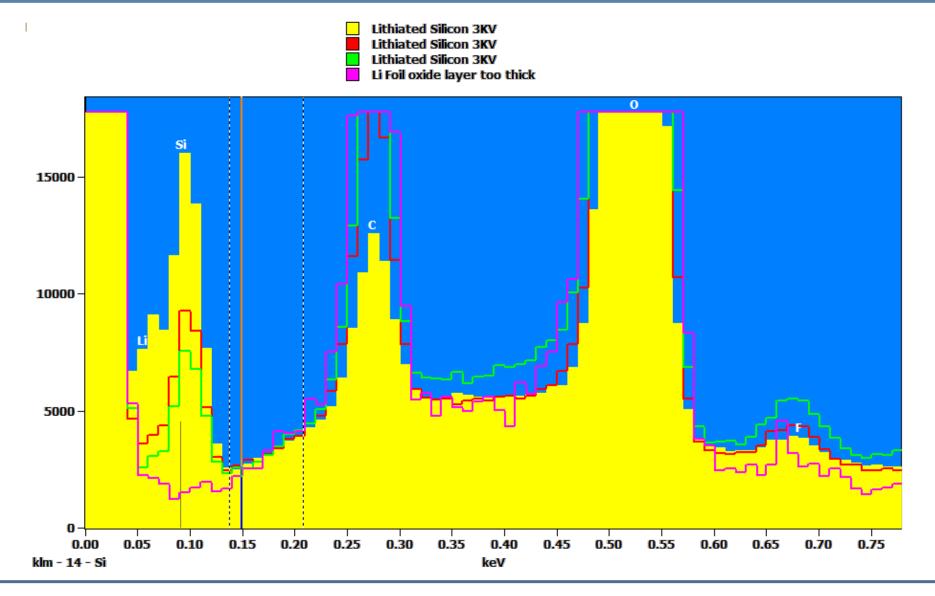


Actual data: Li foil as it oxidizes





Actual data as the Lithiated-Si surface oxidizes





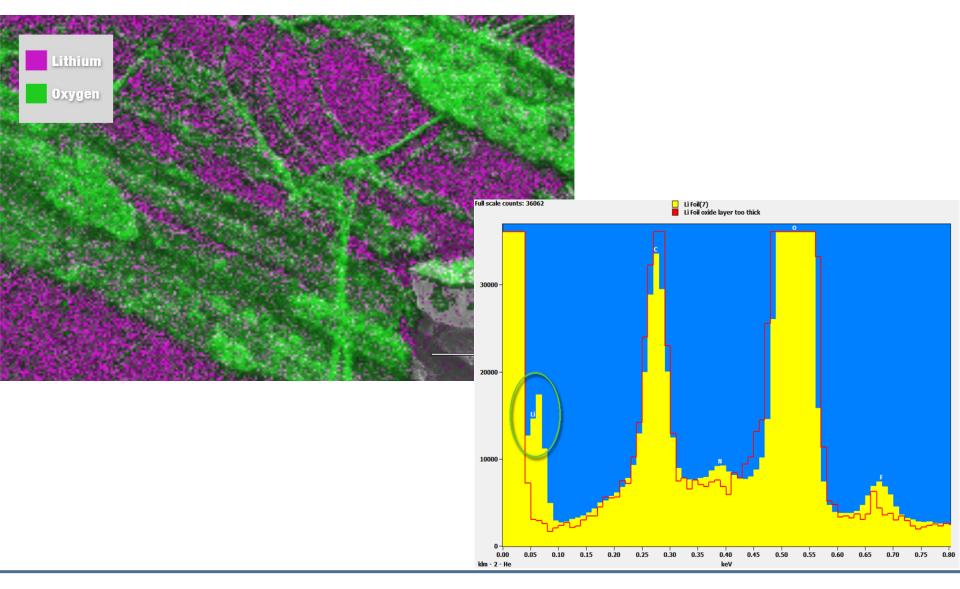


Challenge #3

Mapping Li (and other light elements)

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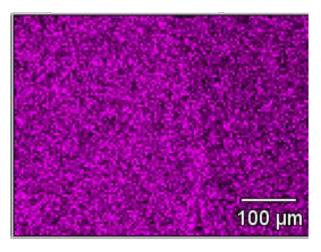
Li mapping



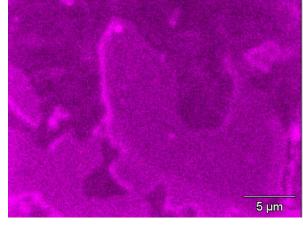


The "real" application of a Li detector

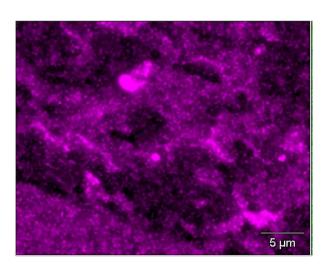
B circa 2008



B circa 2012

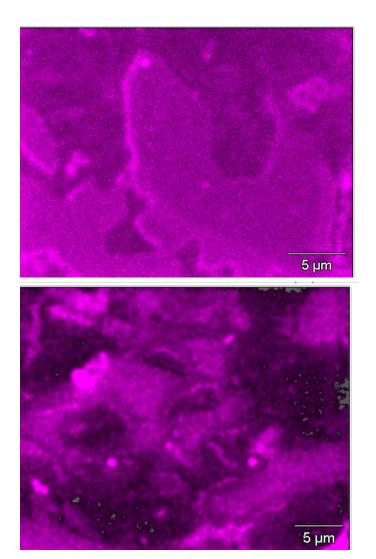


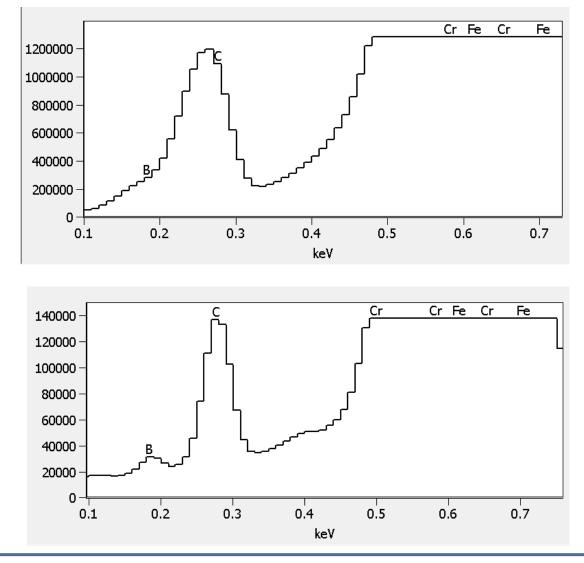
B circa 2014





The "real" application of a Li detector







Summary

- Li detection is possible
- 3 main challenges
 - The detector has specific requirements.
 - The sample must be prepared and maintained very carefully.
 - The sample must be such that the x-rays can escape from the sample and be measured.
- The biggest gain with a Li detector is an overall improvement in light element detection and mapping.

